

## Chapter 18 Chemical Equilibrium Answers Section 3

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### Chapter 18 Chemical Equilibrium Answers

$\text{H}_2(\text{g}) + \text{Br}_2(\text{g}) \rightleftharpoons 2\text{HBr}(\text{g})$   $K = 4.0 \times 10^{-2}$   $K = 3.5 \times 10^4$  Chapter 18 Chemical Equilibrium 18.2 Shifting Equilibrium The Effect of Change in Pressure: Ways to Change Pressure Add or remove a gaseous reactant or product Add an inert gas (one not involved in the reaction) An inert gas increases the total pressure but has no effect on the concentrations or partial pressures of the reactants or products The Effect of Change in Pressure Change the volume of the container When the volume of the ...

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sometimes referred to as the chemical-equilibrium expression. The  $H_2$ ,  $I_2$ ,  $HI$  Equilibrium System Consider the reaction between  $H_2$  and  $I_2$  vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing the rate at which the violet color of the iodine vapor diminishes, as  $[C]^x[D]^y [A]^n[B]^m$  556 CHAPTER 18

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