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Note: Questions with (\*) mark are from syllabus of class XI. Q.1. A solution containing 2.675 g of  $\text{CoCl}_2 \cdot 6\text{NH}_3$  (molar mass =  $267.5 \text{ g mol}^{-1}$ ) is passed through a cation exchanger. The chloride ions obtained in solution were treated with excess of  $\text{AgNO}_3$  to give 4.78 g of  $\text{AgCl}$  (molar mass =  $143.5 \text{ g mol}^{-1}$ ). The formula of the complex is (Atomic mass of  $\text{Ag} = 108 \text{ u}$ )

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